

Supporting Information of  
“Impact of Hydrogen Peroxide on Carbon Corrosion  
in Aqueous KOH Solution”

Atsunori Ikezawa<sup>a,b</sup>, Kohei Miyazaki<sup>a,c,\*</sup>, Tomokazu Fukutsuka<sup>a,c</sup>, and Takeshi Abe<sup>a,c</sup>

<sup>a</sup>*Graduate School of Engineering, Kyoto University, Nishikyo-ku, Kyoto 615-8510, Japan*

<sup>b</sup>*School of Materials and Chemical Technology, Tokyo Institute of Technology, Midori-ku, Yokohama 226-8502, Japan*

<sup>c</sup>*Hall of Global Environmental Research, Kyoto University, Nishikyo-ku, Kyoto 615-8510, Japan*

\* Corresponding Author: (E-mail) [kezawa.a.aa@m.titech.ac.jp](mailto:kezawa.a.aa@m.titech.ac.jp)

(FAX) +81-45-924-5406

Table S1. Electrodeposition conditions of the catalysts to HOPGs

Sample name	Condition	Electrolyte solution (35 cm <sup>3</sup> )
Pt HOPG	Galvanostatic deposition	1 mmol dm <sup>-3</sup> H <sub>2</sub> PtCl <sub>6</sub>
	Current density: -1.80 mA cm <sup>-2</sup> , Time: 3.0 s	+ 2.1 mmol dm <sup>-3</sup> HCl
Ag HOPG	Galvanostatic deposition	2.2 mmol dm <sup>-3</sup> AgNO <sub>3</sub>
	Current density: -1.74 mA cm <sup>-2</sup> , Time: 7.0 s	
MnO <sub>x</sub>  HOPG	Cyclic voltammetry <sup>1,2</sup>	0.1 mol dm <sup>-3</sup> Mn(CH <sub>3</sub> COOH) <sub>2</sub>
	Scan range: 0 – 0.40 V vs. Ag AgCl Sat'd KCl, Scan rate: 20 mV s <sup>-1</sup> , Number of cycles: 5	+ 0.1 mol dm <sup>-3</sup> Na <sub>2</sub> SO <sub>4</sub>
CoO <sub>x</sub>  HOPG	Cyclic voltammetry	0.1 mol dm <sup>-3</sup> Co(CH <sub>3</sub> COOH) <sub>2</sub>
	Scan range: 0 – 0.65 V vs. Ag AgCl Sat'd KCl, Scan rate: 20 mV s <sup>-1</sup> , Number of cycles: 5	+ 0.1 mol dm <sup>-3</sup> Na <sub>2</sub> SO <sub>4</sub>

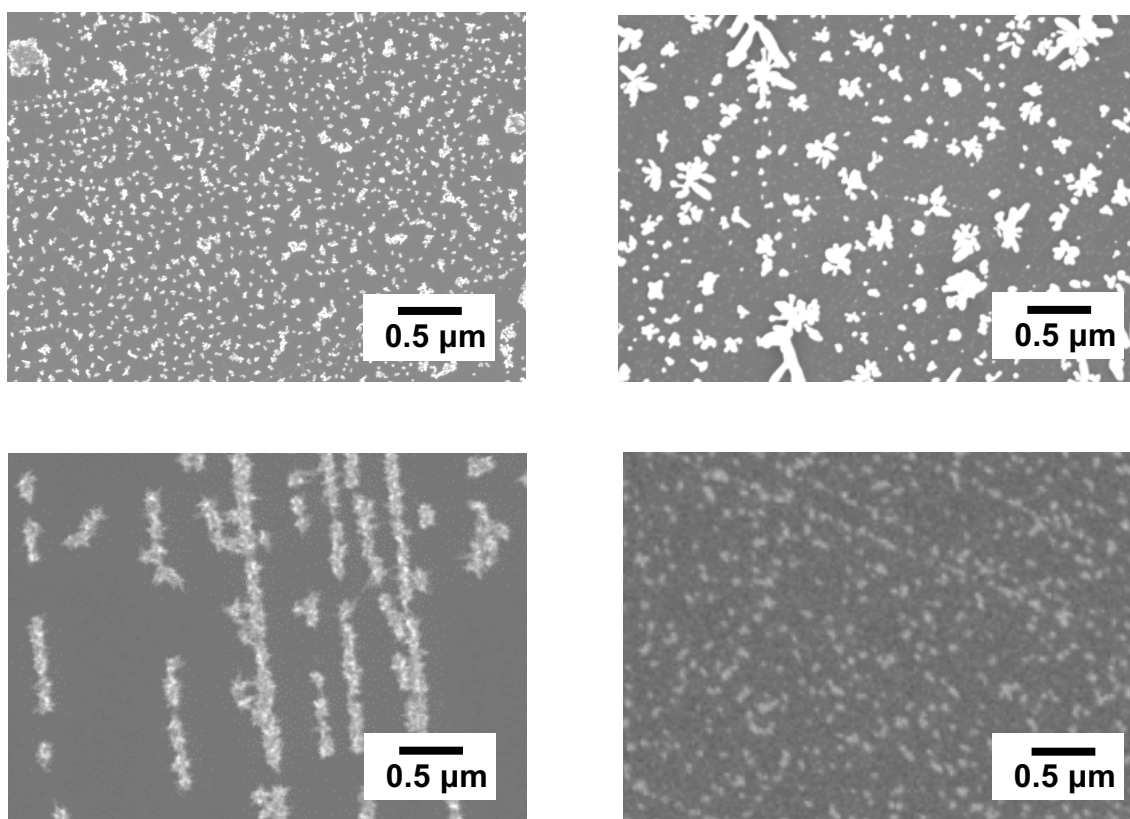


Figure S1. FE-SEM images of the as-prepared (a) Pt|HOPG, (b) Ag|HOPG, (c) MnO<sub>x</sub>|HOPG, and CoO<sub>x</sub>|HOPG.

Table S2. Coverage rates and mean diameters of the catalyst particles on HOPGs analyzed with ImageJ software<sup>3</sup>

Catalyst	Coverage rate / %	Mean diameter / nm
Pt	17	60
Ag	21	130
MnO <sub>x</sub>	14	300
CoO <sub>x</sub>	16	100

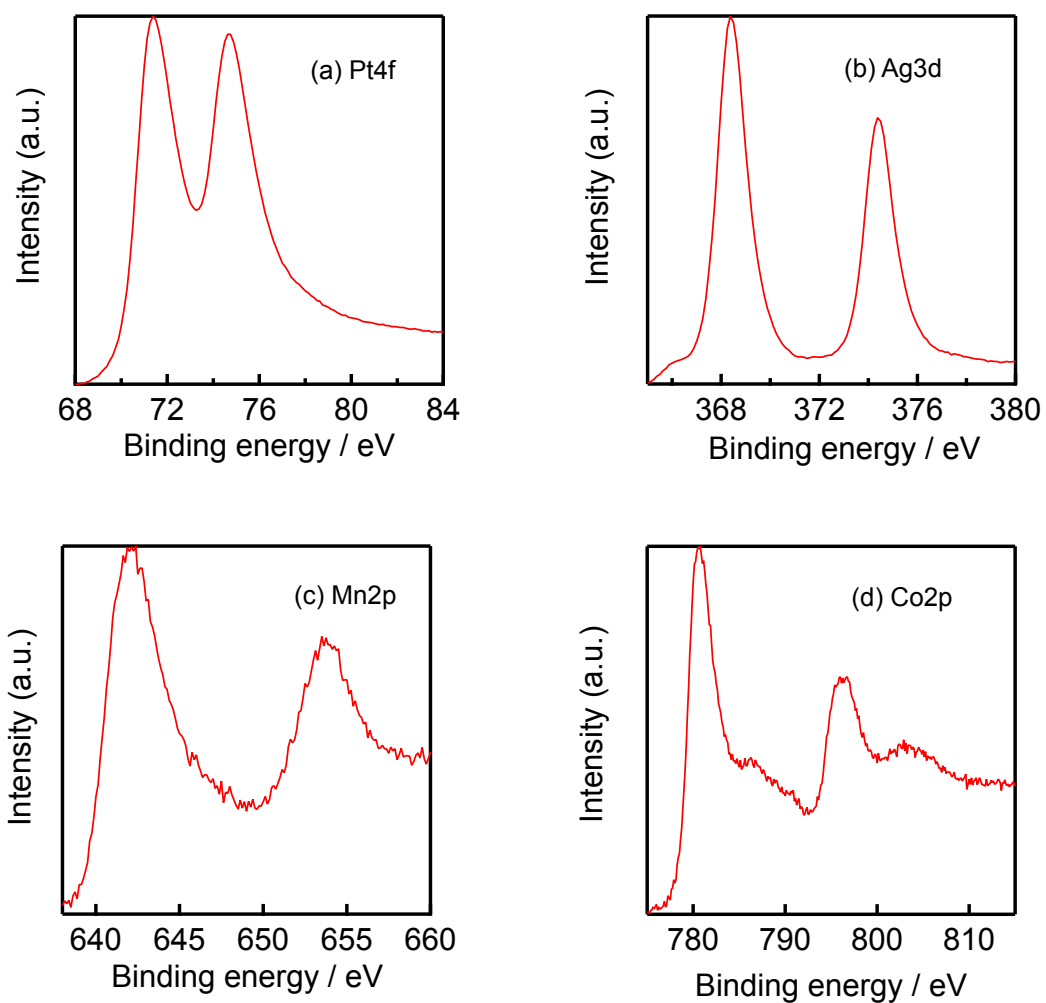


Figure S2. XPS (a) Pt4f, (b) Ag3d, (c) Mn2p, and (d) Co2p spectra of the as-prepared Pt|HOPG, Ag|HOPG, MnO<sub>x</sub>|HOPG, and CoO<sub>x</sub>|HOPG.

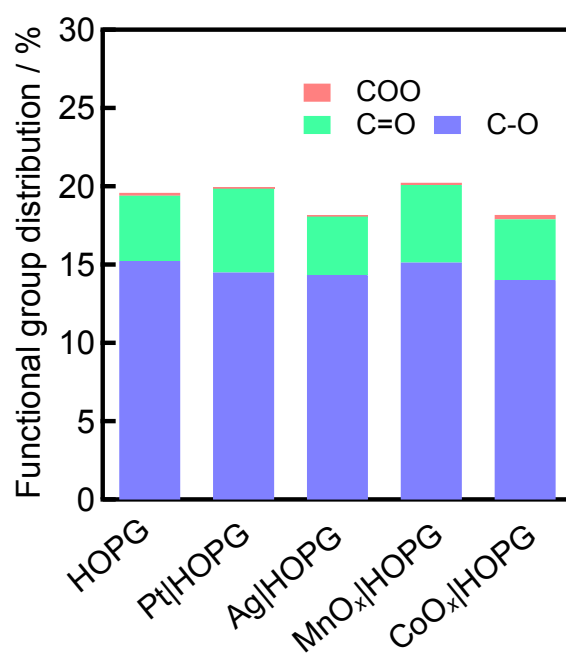


Figure S3. Distribution of oxygen-containing functional groups on the HOPG and catalyst-loaded HOPGs after the immersion test to 1.0 mol dm<sup>-3</sup> KOH + 5 mmol dm<sup>-3</sup> H<sub>2</sub>O<sub>2</sub> calculated from XPS C1s spectra.

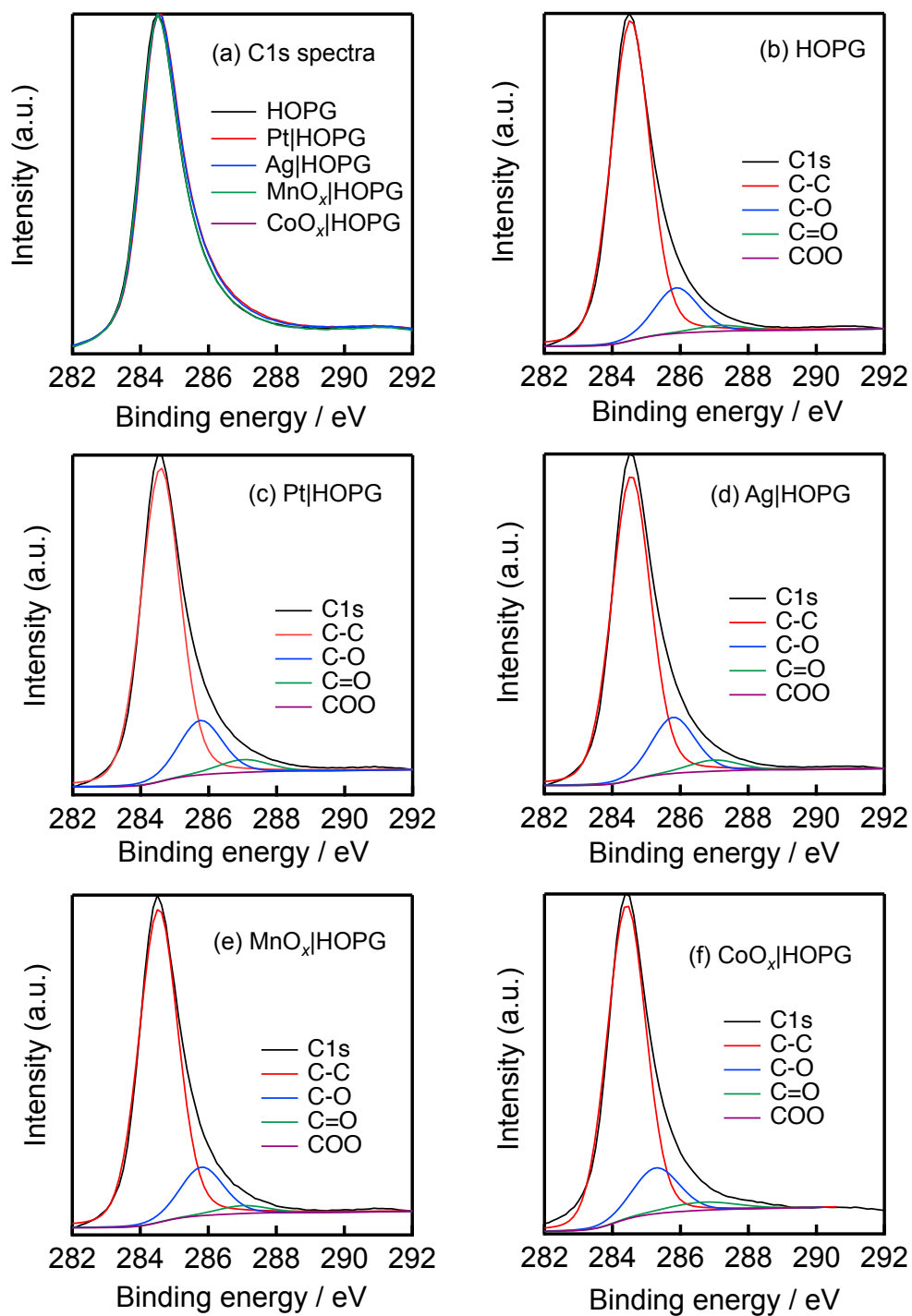


Figure S4. (a) XPS C1s spectra and (b–d) fitting results of the as-prepared HOPG samples.

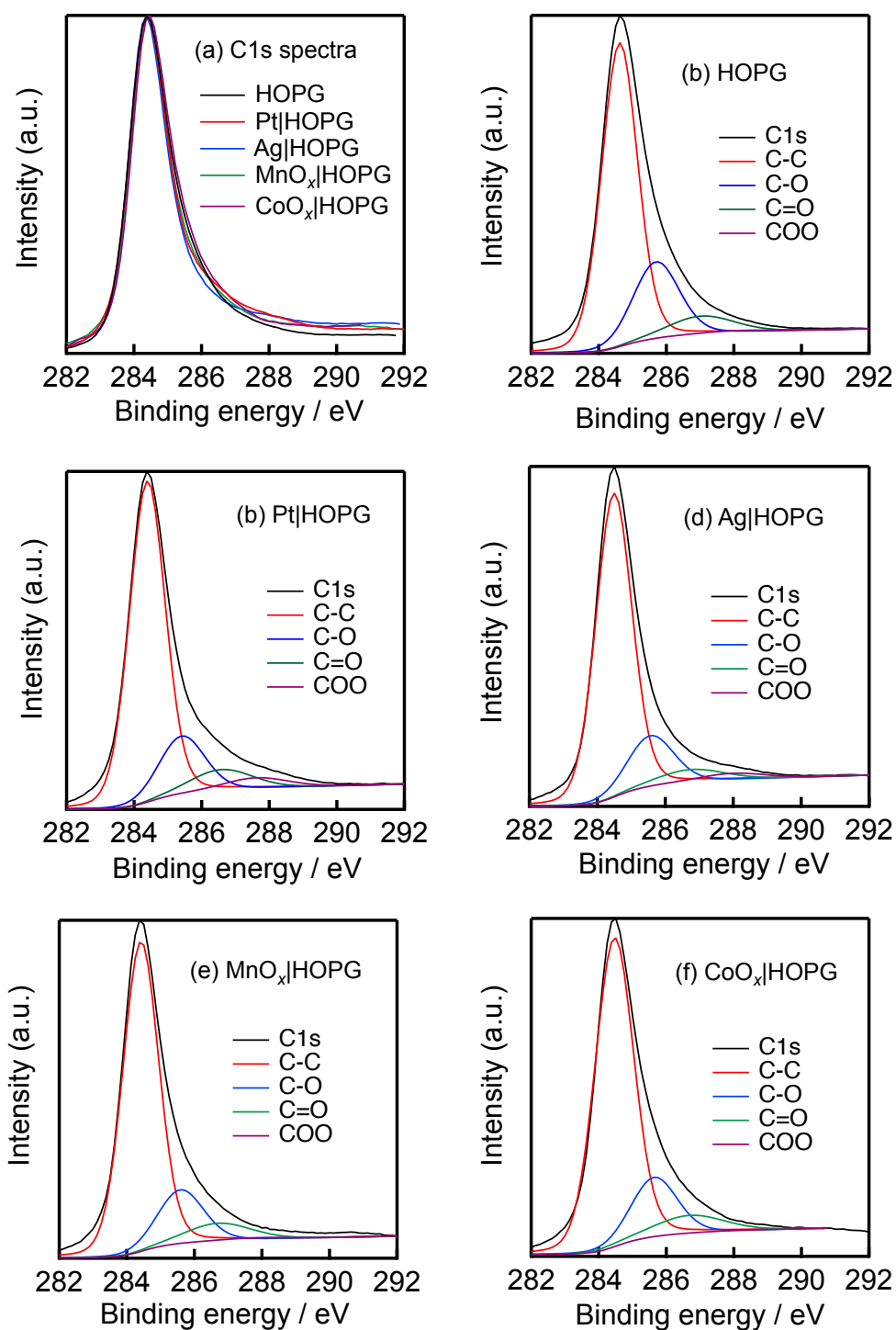


Figure S5. (a) XPS C1s spectra and (b–d) fitting results of the HOPG samples after the immersion test to 1.0 mol dm<sup>-3</sup> KOH + 5 mmol dm<sup>-3</sup> H<sub>2</sub>O<sub>2</sub>.



## References

1. M. S. El-Deab and T. Ohsaka, *Angew. Chem. Int. Ed.*, **45**, 5963 (2006).
2. M. S. El-Deab and T. Ohsaka, *J. Electrochem. Soc.*, **155**, D14 (2008).
3. C. A. Schneider, W. S. Rasband, and K. W. Eliceiri, *Nat. Methods*, **9**, 671 (2012).